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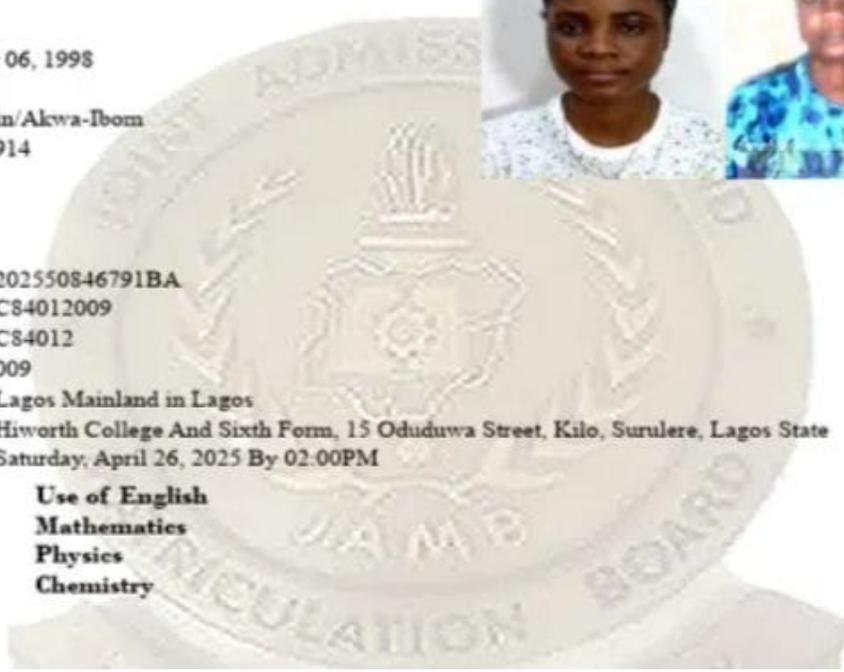


Examination Details:

Registration Number: 202550846791BA
Examination Number: C84012009
Centre Number: C84012
Seat Number: 009
Examination Town: Lagos Mainland in Lagos
Examination Centre: Hiworth College And Sixth Form, 15 Oduduwa Street, Kilo, Surulere, Lagos State
Examination Date: Saturday, April 26, 2025 By 02:00PM

UTME Subjects:
Use of English
Mathematics
Physics
Chemistry

NOTE:



Computer Based Test

USE OF ENGLISH LANGUAGE

1. The smallest particle of an element that retains its chemical properties is:

- A. Atom
- B. Molecule
- C. Ion
- D. Compound

2. The atomic number of an element represents the number of:

- A. Neutrons
- B. Protons
- C. Electrons and neutrons
- D. Protons and neutrons

3. Isotopes of an element have the same:

- A. Mass number
- B. Number of neutrons
- C. Atomic number
- D. Atomic mass

4. Which of the following is a compound?

- A. Oxygen
- B. Hydrogen

- C. Sodium chloride
- D. Iron

5. The formula for methane is:

- A. CH₃
- B. CH₄
- C. C₂H₆
- D. CO₂

6. The process by which a solid changes directly to gas is called:

- A. Evaporation
- B. Condensation
- C. Sublimation
- D. Melting

7. Which of the following increases the rate of evaporation?

- A. Low temperature
- B. High humidity
- C. Large surface area
- D. High pressure

8. The kinetic theory states that particles in a gas move:

- A. Slowly
- B. Randomly
- C. In circles
- D. Not at all

9. The boiling point of a liquid depends on:

- A. External pressure
- B. Colour
- C. Shape
- D. Size

10. Diffusion is fastest in:

- A. Solids
- B. Liquids
- C. Gases
- D. Plasma

11. Elements in the same group have the same number of:

- A. Shells
- B. Valence electrons
- C. Neutrons
- D. Protons

12. The most reactive group of metals is:

- A. Group I
- B. Group II
- C. Group VII
- D. Group 0

13. Chlorine belongs to which group?

- A. Group I
- B. Group VII
- C. Group II
- D. Group 0

14. Noble gases are chemically:

- A. Very reactive
- B. Reactive
- C. Inert
- D. Metallic

15. Atomic size decreases across a period because:

- A. Nuclear charge increases
- B. Shielding increases
- C. Mass increases
- D. Protons decrease

16. Ionic bonding involves transfer of:

- A. Protons
- B. Neutrons
- C. Electrons
- D. Atoms

17. Covalent bonding involves:

- A. Sharing of electrons
- B. Transfer of protons
- C. Loss of neutrons
- D. Breaking of atoms

18. Which compound is ionic?

- A. CO_2
- B. HCl
- C. NaCl
- D. NH_3

19. A double bond consists of how many shared electron pairs?

- A. One
- B. Two

- C. Three
- D. Four

20. The shape of methane is:

- A. Linear
- B. Trigonal planar
- C. Tetrahedral
- D. Octahedral

21. One mole of a substance contains:

- A. 6.02×10^{23} particles
- B. 6.02×10^{22} particles
- C. 6.02×10^{21} particles
- D. 6.02×10^{20} particles

22. The molar mass of H_2O is:

- A. 16 g/mol
- B. 17 g/mol
- C. 18 g/mol
- D. 20 g/mol

23. Relative atomic mass of carbon is:

- A. 10
- B. 12
- C. 14
- D. 16

24. The empirical formula of C_2H_4 is:

- A. CH
- B. CH_2
- C. C_2H_2
- D. C_2H_4

25. At STP, one mole of gas occupies:

- A. 22.4 dm^3
- B. 24.0 dm^3
- C. 20.0 dm^3
- D. 18.0 dm^3

26. A substance that produces H^+ in solution is a:

- A. Base
- B. Salt
- C. Acid
- D. Metal

27. The pH of a neutral solution is:

- A. 1
- B. 7
- C. 14
- D. 10

28. Which of the following is a base?

- A. HCl
- B. NaOH
- C. H₂SO₄
- D. HNO₃

29. Blue litmus turns red in:

- A. Base
- B. Neutral solution
- C. Acid
- D. Salt

30. A salt is formed from:

- A. Metal + oxygen
- B. Acid + base
- C. Gas + gas
- D. Metal only

31. Oxidation involves:

- A. Gain of electrons
- B. Loss of electrons
- C. Gain of protons
- D. Loss of neutrons

32. Reduction involves:

- A. Loss of electrons
- B. Gain of electrons
- C. Loss of oxygen
- D. Gain of hydrogen only

33. An electrolyte conducts electricity because it contains:

- A. Molecules
- B. Atoms
- C. Ions
- D. Neutrons

34. Rusting of iron requires:

- A. Oxygen only
- B. Water only

- C. Oxygen and water
- D. Nitrogen

35. The positive electrode in electrolysis is the:

- A. Cathode
- B. Anode
- C. Cell
- D. Ion

36. Alkanes are:

- A. Saturated hydrocarbons
- B. Unsaturated hydrocarbons
- C. Aromatic compounds
- D. Alcohols

37. The general formula of alkanes is:

- A. C_nH_{2n}
- B. C_nH_{2n+2}
- C. C_nH_{2n-2}
- D. C_nH_n

38. Ethene belongs to the:

- A. Alkane family
- B. Alkene family
- C. Alkyne family
- D. Alcohol family

39. Fermentation produces:

- A. Methane
- B. Ethanol
- C. Propane
- D. Butane

40. Benzene is an example of a:

- A. Saturated compound
- B. Aromatic compound
- C. Alkane
- D. Alcohol

41. Hard water contains high amounts of:

- A. Sodium ions
- B. Calcium ions
- C. Chloride ions
- D. Hydrogen ions

42. The process of removing temporary hardness is called:

- A. Filtration
- B. Distillation
- C. Boiling
- D. Crystallization

43. A catalyst works by:

- A. Increasing activation energy
- B. Decreasing activation energy
- C. Changing products
- D. Being consumed

44. Haber process is used to manufacture:

- A. Ammonia
- B. Sulphuric acid
- C. Nitric acid
- D. Sodium hydroxide

45. Contact process produces:

- A. NH_3
- B. H_2SO_4
- C. HNO_3
- D. NaCl

46. Radioactivity was discovered by:

- A. Dalton
- B. Rutherford
- C. Becquerel
- D. Bohr

47. Alpha particles carry a charge of:

- A. +1
- B. +2
- C. -1
- D. 0

48. Beta particles are:

- A. Helium nuclei
- B. Electrons
- C. Protons
- D. Neutrons

49. Gamma rays are:

- A. Particles
- B. Electromagnetic waves

- C. Protons
- D. Electrons

50. Half-life is the time required for:

- A. Complete decay
- B. Half of a substance to decay
- C. Full reaction
- D. Boiling